

# DETERMINATION OF TOTAL ALKALOID CONTENT IN CORTEX CHINAE (TITRIMETRIC)

*PHARMACOGNOSY-III PRACTICE*  
2020

## Methods for the determination of active substances in drugs:

- Gravimetric
- Titrimetric
- Volumetric
- Colorimetric
- Chromatographic (HPLC, GC)
- Spectrophotometric
- Polarimetric

Plant Name: *Cinchona succirubra* (Rubiaceae)

Drug Name: Cortex Chinae

**Alkaloid mixture** (quinine-quinidine, cinchonine-cinchonidine)

The total alkaloid content is estimated as  
quinine and cinchonine.

The alkaloids generally exist in the plant as salts of acids and sometimes as bound with tannins.

Alkaloid extraction is carried out by two methods:

- 1) Using water or aqueous alcohol containing dilute acid
- 2) Using organic solvent in alkaline medium

# Experimental procedure

Quinic and quinotannic acid

25% formic acid (2 ml)  
+water (15 ml)

extract in the water bath for 30 minutes with continuous stirring using erlenmeyer flask with cap

↓ cool

+ chloroform (15 ml)

+ ether (60 ml)

+30% NaOH (5 ml)

↓ Shake for 1 min

+ Gum tragacanth(2 g)

↓ Shake vigorously for 1 min

take the ether phase at the top

The ether phase is evaporated in distillation apparatus until 1-2 ml remains.

Residue

+ ethanol (10 ml)  
+ boiled water (10 ml)  
+ methyl red (10 drops)

Titration up to red colour with 0.1 N HCl  
(A ml 0.1 N HCl)

Methyl red is a pH indicator;  
it is red in pH under 4.4,  
yellow in pH over 6.2

# CALCULATION

1l	1N HCl	36.5 g
1l	0.1N HCl	3.65 g
1ml	0.1N HCl	<b>0.00365 g</b>

$$\begin{aligned} \text{MA}_{\text{quinine} = \text{quinidine}} &= 324.40 \text{ g} \\ \text{MA}_{\text{chinchonine} = \text{chinchonidine}} &= + 294.40 \text{ g} \quad /2 \\ \hline \text{Accepted as average molecular} & \\ \text{weight for alkaloids in the drug} & \quad \quad \quad \mathbf{309.4 \text{ g}} \end{aligned}$$

1 molecule HCl reacts with 1 molecule of alkaloid  
 36.5 g of HCl reacts with 309.4 g of alkaloid  
 0.00365 g of HCl reacts 30.94 mg of alkaloid

1 ml 0.1 N HCl is equivalent to 30.94 mg of alkaloid  
A ml 0.1 N HCl is equivalent to X mg of alkaloid  
 $X \text{ mg} = X \cdot 10^{-3} \text{ g}$

1,25.. g drog X.10<sup>-3</sup> g of alkaloid  
100 g drog B g of alkaloid

B = % alkaloid content