

CEN4415 PROCESS DESIGN I

2.2 Predicting Trends

In this method, there are four possible modifiers to a term in an equation.

- 1. ¢: value remains constant.
- 2. \uparrow : value increases.
- 3. \downarrow : value decreases.
- 4. ?: value change not known.

Each variable in an expression has one of these identifiers appended. If x is a constant then $x \rightarrow x(c)$.

Example

For a bimolecular, elementary, gas-phase reaction, the rate expression is known to be $-r_A = kC_AC_B$.

a. What is the effect on the reaction rate of increasing the reaction pressure by 10% while maintaining constant temperature?

b. What is the effect on the reaction rate of increasing the reaction temperature by 10% while maintaining constant pressure?

SOLUTION :

The goal is to determine the effect of increasing pressure on the reaction rate. The equation containing the unknown value, which is the reaction rate, rA, is given. We need a relationship between pressure and at least one variable in the equation.

We have information, from the ideal gas law, that tells us that

Ci = Pi/RT,

where i is A or B.

Now that we have the necessary relationships, an intuitive understanding can be obtained by predicting trends.

a)

 $Ci(\uparrow) = yi(c)P(\uparrow)/R(c)T(c) = Pi(\uparrow)/R(c)T(c)$

Moving to the rate expression, the resulting trend is predicted by

 $rA(\uparrow) = k(c)CA(\uparrow)CB(\uparrow)$

which means that the reaction rate increases.

b)

 $Ci(\downarrow) = yi(c)P(c)/R(c)T(\uparrow) = Pi(c)/R(c)T(\uparrow)$

From the rate expression,

 $rA(?) = k(\uparrow)CA(\downarrow)CB(\downarrow)$

REFERENCES

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Chemical Engineering Design, ButterWorth Heinemann, Oxford.

2. Turton R., Bailie R.C., Whitin W.C., Shaeiwitz J.A. 1998, Analysis, Synthesis and Design of Chemical Processes, Prentice Hall, New Jersey.